

02. Mini-Review: Syngas Production Via Partial Oxidation of Methane Reaction and Its Potential Catalyst

by Silvana Dwi Nurherdiana

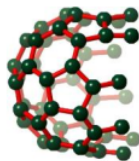
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Mini-Review: Syngas Production Via Partial Oxidation of Methane Reaction and Its Potential Catalyst

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Abstract

Methane as a light gas was generally found in natural gas, which was burned freely to gain a high quality of petroleum. This action truly impacted the worst condition in nature, namely the greenhouse effect. This brief review described a fundamental theory of the crucial process in methane conversion from natural gas into value-added chemicals such as syngas (CO+H₂). The methane conversion reaction was commonly divided into direct and indirect reactions. The indirect reaction such as partial oxidation of methane was mostly chosen due to the intermediate product (syngas) can easily generate many raw materials of petrochemicals. This paper also described a potential catalyst to be applied in heterogeneous types, such as perovskite oxide, metal oxide, and zeolite.

Keywords: methane; syngas; perovskite oxide; catalyst, methane conversion.

Abstrak

Metana sebagai gas ringan umumnya terdapat dalam gas alam yang dibakar secara bebas untuk mendapatkan minyak bumi yang berkualitas tinggi. Tindakan ini benar-benar berdampak pada kondisi alam yang paling buruk yaitu efek rumah kaca. Tinjauan singkat ini menjelaskan teori dasar tentang proses penting dalam konversi metana dari gas alam menjadi bahan kimia bernilai tambah seperti syngas (CO+H₂). Reaksi konversi metana secara umum dibagi menjadi reaksi langsung dan tidak langsung. Reaksi tidak langsung seperti oksidasi parsial metana banyak dipilih karena produk antara (syngas) dapat menghasilkan banyak bahan baku petrokimia. Makalah ini juga memaparkan katalis yang berpotensi untuk diaplikasikan yaitu tipe heterogen seperti oksida perovskit, oksida logam dan zeolit.

Kata kunci: metana, syngas, oksida perovskite, katalis, konversi metana.

Introduction

Indonesia has abundant natural gas reserves which methane as the main component with the 11th largest reserves rank in the world and a reserve to production value (R/P) of 59 years at a stable price [1]. In addition, natural gas also accounts for about 20% of long-chain hydrocarbons and non-hydrocarbon gases such as N₂, CO₂ and H₂S [2]. The highly fuel content in natural gas makes it widely consumed as a source of energy for industry and households.

Although the reserves are in abundant amounts, the use of natural gas is limited by the distribution constraint due to its form is gas. The long-chain hydrocarbon gas

component can be compressed into a liquid by applying high pressure known as CNG (Compressed Natural Gas), which makes it easier to control and move. However, its greatest contain is methane, is a gas that can not be compressed so that its use really can only be done in the form of gas. Transport constraints lead to the most effective and efficient utilization only when close to natural gas sources. Based on the description, the increase in the benefits of natural gas will be expanded if methane in natural gas is converted into longer chain hydrocarbons that can be melted and become more value-added petrochemical materials such as polymer, fertilizer and plastic raw materials [3].

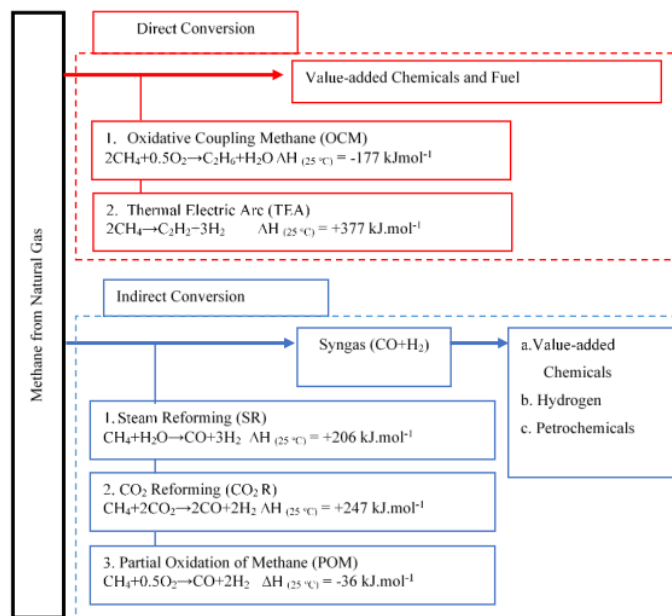
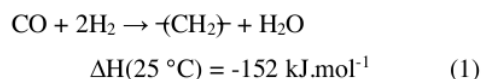


Figure 1. The reaction of Methane Conversion from Natural Gas

Conversion of methane into long-chain hydrocarbons can be converted through the direct and indirect processes as shown in Figure 1. Direct conversion methods such as Oxidative Coupling Methane (OCM) have been studied by Farsi and Mansouri [4] with methane and yield (C₂₊) conversion respectively reaching 45% and 27%. This value can be achieved in reaction with high temperature and low reaction efficiency kinetically. Another inefficient method is the Thermal Electric Arc (TEA) which requires operation temperatures greater than 2727 °C from plasma with a power of 10 kWh.kg⁻¹-C₂H₂ [5].

Compared to the direct conversion methods, indirect conversions such as Steam Reforming (SR) and CO₂ Reforming (CO₂ R) react in an endothermic condition that requires enormous energy and expensive maintenance costs. Indirect conversion is performed by producing an intermediate product of syngas (mixture of H₂ and CO with ratio 2:1), followed by long-chain hydrocarbon synthesis using the Fischer-Tropsch method with the reaction shown by Eq 1. However, the ratio of H₂:CO generated from the methods of SR and CO₂ R sequentially is 3: 1 and 1: 1, which do not correspond to the amount of feed required for the Fisher-Tropsch synthesis of 2: 1.



Partial oxidation of methane (POM) also produces syngas with a stoichiometric H₂:CO of 2:1 which ideal ratio for hydrocarbon production by the FT process and react in an exothermic condition which can reduce 30% operating costs. In addition, POM reaction can also be performed on a large scale (industrial scale) or small scale (laboratory scale) [6]. The syngas usage for the industry have been much attention as a raw material for methanol industry such as Methanol Bunyu Pertamina and PT Kaltim Methanol Industry with the total production capacity of 990,000 tons per year, PT Sintas Kurama Perdana 11,000 tons per year and imports 3,532,323 tons in 2004 and ammonia industries such as PT Pupuk Kaltim, Kaltim Pacific Ammonia, and PT Kaltim Parna with total production capacity 3,860,000 tons per year.

Therefore, the development and optimization of the syngas production process become very important to do [7]. The POM reaction, although it has been done with both homogeneous and heterogeneous catalysts which requires a high temperature of 700 to 1,000 °C to activate four identical C-H bonds. In addition, the oxygen involved in the reaction should not be excessive due to high temperatures and oxygen in excess amounts lead to the complete oxidation of methane into CO₂ dan H₂O. Therefore, this article briefly described the fundamental overview of syngas, partial oxidation of methane reaction

and potential catalyst which involved in optimizing the syngas production

Fundamental Theory of Syngas

Syngas is an intermediate product of methane conversion reaction, composed of H₂ and CO with a composition ratio of 2:1 [8]. Hereinafter, Table 1 describes the following characteristics of H₂ and CO. Moreover, chemical characteristics can also be known from the molecular structure of H₂ and CO. Hydrogen is a non-metallic element, colorless and odorless, binding to form covalent diatomic hydrogen due to the use of electrons together, highly flammable but not emitting pollutants into the environment at standard temperature and pressure. Therefore, hydrogen as the most considerable energy has energy per unit of weight of 120,7 kJ.g⁻¹ thus recommended as carrier energy in the future [9].

Table 1. H₂ and CO Gas Characteristics

Characteristics	H _{2(g)}	CO _(g)
Density (g.L ⁻¹) (0 °C, 1 atm)	0,08988	1,250
Molar mass (g.mol ⁻¹)	2	28,0101
Melting point (°C)	-259,14	-205
Boiling point (°C)	-252,87	-192

Source : Harahap and Tjahjono, 2016 [7]

While CO has chemical bonds consisting of two covalent bonds and a

coordinate covalent bond with the electron density at C is higher than O which introduces a small dipole moment with a small electronegative difference. As a result, the molecular structure can resonate then affects the reactivity to other carbon cations. CO is colorless, odorless that can produce from incomplete combustion of carbon or methane compound. Despite its toxic and flammable nature, CO plays a key role as a primary feedstock for chemicals products such as methanol, acetic acid and long-chain hydrocarbon fuels from hydrogenation and the Fischer-tropsch reactions between CO and H₂ [10].

In addition, syngas is also used as fuel gas and produce value-added chemicals as shown in Figure 2 which can be utilized by the petrochemical industry for the raw materials of polymers, textile, fertilizers and farms. Syngas can be obtained basically from the methane conversion process from natural gas and coal. While coal as a raw material produces harmful and toxic emissions during the syngas production process. Thus natural gas has been widely consumed to reduce the risk of global warming. Indonesia is the country with the largest natural gas reserves at 11th place from all over the world and 3rd from Asia-pacific with 99.77 trillion cubic feet (TCF) with a reserve potential of 50.21 TCF [11]. Natural gas sources can be obtained from biomass gasification (biogas) using

anaerobic bacteria by purification process of water (H_2O), hydrogen sulfide (H_2S) and particulate to reach pipeline quality. Besides natural gas from biogas, other sources of natural gas are oil, coal and natural gas fields.

Both sources have the largest methane composition with a number of differences in secondary compositions as shown in Table 2 [12].

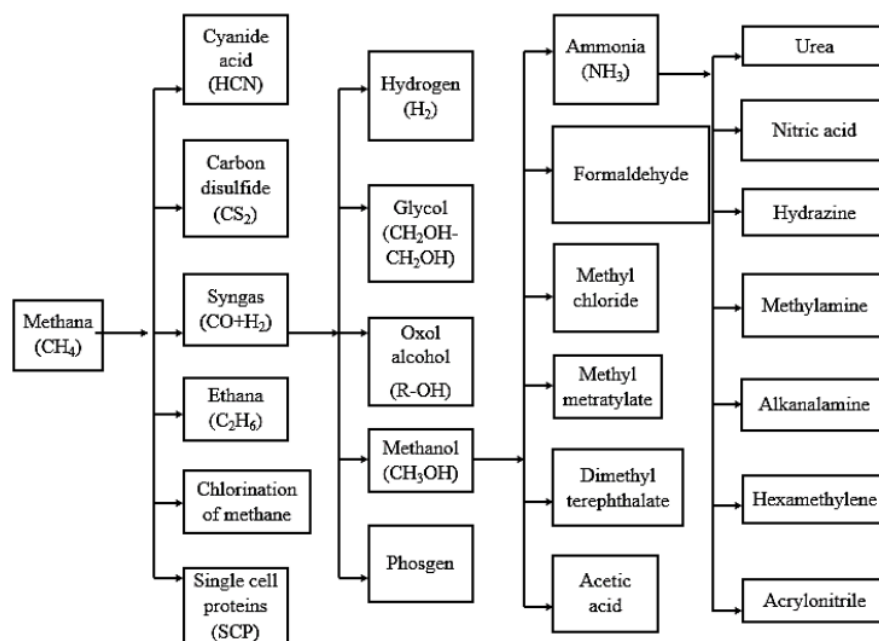


Figure 2. Methane Conversion Products [10].

Meanwhile, the high methane composition in nature has a serious impact in radiation up to 21 times greater than carbon dioxide (CO_2), greenhouse effect and rising earth temperature. This is because methane is very easy to react photochemically with OH radicals derived from the reaction between

water vapor and ozone (O_3) in the troposphere so that the amount of O_3 decreases [13]. Therefore, natural gas in mining areas contributes the most significant percentage of methane content which is preferably burned to carbon dioxide.

Table 2. Biogas and Natural Gas Composition

Compound	Composition (%)	
	Biogas	Natural gas
Methana (CH ₄)	55-70	60-90%
Carbon dioxide (CO ₂)	30-40	0,61
Nitrogen (N ₂)	0-0.3	1
Hydrogen sulfide (H ₂ S)	0-3	0.07
Oxygen (O ₂)	0.1-0.5	-
Hydrogen (H ₂)	1-5	-
Ethane (C ₂ H ₆)	-	5-8
Propane (C ₃ H ₈)	-	2-7
Butane (C ₄ H ₁₀)	-	0,4-3
Pentane (C ₅ H ₁₂)	-	0,2-5

On the other hand, natural gas can be utilized as household fuel for large-scale purposes such as industry rather than burned in vain. However, the usage can be utilized directly and close to natural gas sources to maintain the exergy efficiency. However, most of the natural gas reserves are scattered

in the offshore areas as seen in Figure 3 of the Natuna, South Sumatra, East Kalimantan, Masela and West Papua which areas are far from residential areas thus requiring consideration of production and maintenance techniques as well as the investment costs for gas-distributing such as a gas pipe.

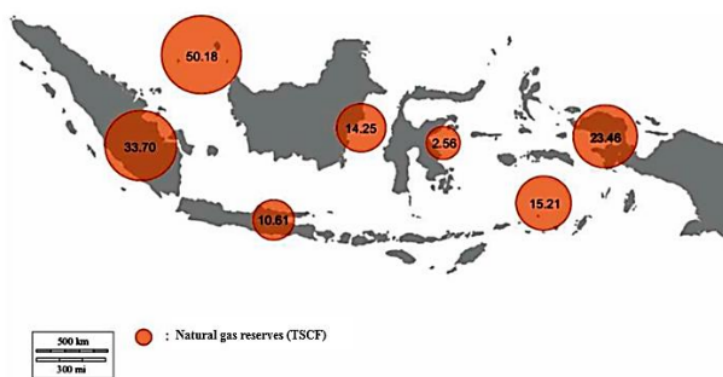


Figure 3. Natural Gas Reserves in Indonesia [11]

In addition, natural gas reserves in Indonesia increased by 1.5% from 2004 to 2015 and estimated the reserve values compared to the production of 37.8 [14]. The value-added products from natural gas need to be increased such as chemicals to balance the operating costs of the energy industry. Thus, a process of methane conversion to syngas is essential to be optimized.

Methane Conversion To Syngas Via Partial Oxidation of Methane (POM) Reaction.

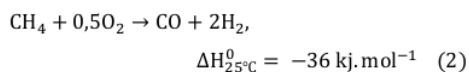
Syngas can be obtained through a methane conversion reaction. In order to the methane conversion process by direct and

indirect conversion as mentioned above, the process has each advantage and disadvantages in terms of the operation process and operational costs. As described in Table 3, the syngas production process using POM reaction is mainly recommended to other conventional processes. In addition, POM reaction can be applied on a small scale (laboratory) to large scale such as industrial production to be commercialized [15]. POM reaction occurs exothermic and irreversible which produce 36 kJ per mole of heat thus the reaction is being considered due to high efficiency of exergy.

Table 3. Methane Conversion Conditions and Results

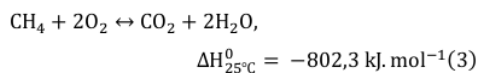
Conversion methods	Temperature (°C)	ΔH (25 °C) (kJ.mol ⁻¹)	Performance	Ref
OCM	650-950	-177	C ₂ Selectivity 70-90%	[16]
TEA	727-2727	+377	Acetylene Selectivity 70-80%	[5]
SR	625-750	+206	Ratio of H ₂ /CO = 2 - 3	[17]
CO ₂ R	700	+247	High resistancy on carbon deposition	[18]
POM	600-950	-36	CH ₄ Conversion: 98 % dan CO Selctivity: 97.5 %	[19]

POM reaction is imperfection combustion due to involving only half of the amount of oxygen by the reaction in Eq. 2 as follows.

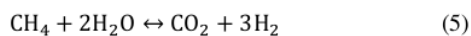
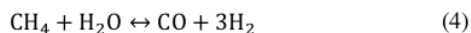


Another advantage of the POM reaction is the lower reactor heat which consumes lower operational cost and H₂:CO ratio production of about 2 which can be used in the Fischer-Tropsch process. However, a combustion reaction also occurs when methane reacts with

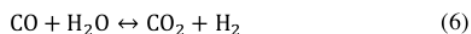
oxygen because lower exothermic value as shown in Eq. 3 [20].



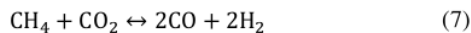
Furthermore, excess methane will react with vapor endothermically through the steam reforming process. This combination of POM and SR reaction will make the process energy more efficient by the reaction shown in the following equations [2] :



Followed by water gas shift reaction:



And CO₂ R reaction to form syngas:



The selectivity of syngas can be reduced due to the carbon deposition formation which increases the deactivation of the catalytic site of the catalyst by the following Boudouard reaction:



The carbon fraction can be gasified by heating or oxygen by the reaction:



Therefore, the major product of syngas can be obtained by particularly control of the O₂:CH₄ ratio as feedstock and the temperature conditions more than 1000 °C with a very short thermodynamic contact time. In addition, the catalyst can be used to control syngas production such as metal oxides, zeolites and perovskite oxides [21].

Active Catalyst Systems for Pom Reaction

As mentioned before, the catalyst plays an important role in the exergy efficiency of POM reaction and optimization of the methane conversion process as well as the selectivity of syngas. Some mostly recommended catalysts are metal oxides, zeolites and perovskite oxides which have different catalytic performances as shown in Table 4

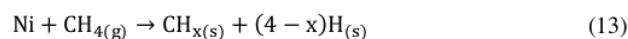
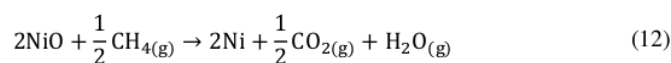
The metal from metal oxides catalysts is derived from two types: noble metal catalyst (Rh, Ru, Pt, Pd, Ir) and non-noble metal catalysts (Ag, Pt, Ni, Co) [25]. The noble metal catalyst has higher catalytic activity and is hardly poisoned by carbon deposition compared to a non-noble metal catalyst. According to an economical subject, non-noble metals are more economic and widely recommended in the industry [27]

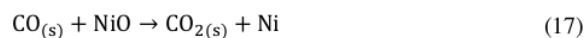
Table 4. Metal Oxide, Zeolite and Perovskite Catalyst for Partial Oxidation of Methane

Catalyst	Temperature (°C)	Time (h)	CH ₄ Conversion (%)	Syngas		Ref
				Selectivity (%)		
				H ₂	CO	
Ni/Al ₂ O ₃	900	-	80	-	90	[22]
Ni/MgO	750	100	91.2	96.3	92.4	[23]
Ni-Mg/Al ₂ O ₃	800	-	96.5	94	95	[24]
NiO(75)-Al ₂ O ₃	650	25	65	80	70	[25]
Ni/zeolit Y	700	89.6	-	88.7		[26]
Mesopori NiO-Al ₂ O ₃	600	40	80	76	95	[27]
Ni-zeolit BEA	625-750	4-8	6-7	-	40- 50	[17]
Pt/10%Rh	800	8	28.2		94	[28]
Pt-CeO ₂	800	-	98	61	48	[29]
Ag/CGA	600	-	23.4	-	94	[30]
La _{0.6} Ca _{0.4} Fe _{0.75} Co _{0.25} O _{3-δ}	900	1400	99	-	83	[31]
La _{0.6} Sr _{0.4} Ti _{0.3} Fe _{0.9} O _{3-δ}	900	10	60	-	99.9	[32], [33]

Nickel catalyst has high catalytic activity and resistance to carbon deposition due to the presence of supporting material to reduce the potential for deactivation as well as membrane leakage of high operational temperature. Mostly, nickel catalyst is easily obtained from NiO as metal oxide or nitrate compound

which has high electron conductivity and can increase the CO and H₂ selectivity. As reported by Au et al that NiO/Al₂O₃ was successfully used for syngas production at 500 °C with a recommendation reaction as follows [34].





Compared to the reaction in equation (13), reaction (12) is more difficult to react thus a pre-treatment process is essentially required to activate the catalyst active site such as flowing H_2 at 500 °C for an hour to reduce metal oxide into metal or the other treatment usage of carrying out the POM reaction at a temperature above 800 °C. This is because the metal oxide catalyst is only active in oxidation number of 0 then activates the CH_4 to produce CO , H_2 and CO_2 . The characteristics also apply to other metal oxide catalysts which are activated only in a metal form such as Rh, Pt and Co [[34], [35]].

In addition to metal oxide, the zeolite may also act as a POM reaction catalyst and good catalyst support simultaneously since the nickel catalyst is less dispersed on a metal oxide support. Zeolite has been widely applied as a heterogeneous catalyst because it has a regular pore structure, high surface area and acidic properties. These advantages make zeolites have a higher methane conversion of 94% and H_2 selectivity of 93% than metal oxide as a catalyst supporter [[26]]. As stated in some previous research, suitable zeolite can

be used as a POM reaction catalyst. They are ZSM5, Zeolite Y, mordenite and BEA zeolite. The catalysts have a surface area between 400 to 800 $\text{m}^2 \cdot \text{g}^{-1}$ which Ni easily dispersed well. Moreover, the acid side increased the toxicity resistance such as CO_2 that able to deactivate the catalyst [36].

The catalyst usage can be adjusted to decrease the reactor heating temperature, carbon deposition as well as can control the feed for POM reaction to syngas. The use of environmental air as an oxygen supply can improve efficiency because the use of pure oxygen operationally increases the risk of fire and excessive heat which also increases operational cost. However, the content of 70% of the air is nitrogen while the separation between nitrogen and oxygen is not easy [[20]]. In 1985 Tereoka et al introduced perovskite as an ionic and electronic conductor that continues to be applied as a solid oxide fuel cell (SOFC) to a methane conversion catalyst such as POM reaction. As aforementioned, oxygen does not diffuse through the pore but rather the oxygen lattice in the perovskite structure [37].

Currently, heterogeneous catalysts are widely studied in the form of metal oxides, zeolites and perovskites. Metal oxides used include Al_2O_3 , NiO , CeO_2 dan some other transition metal oxides [[26], [38], [39]]. Some researchers suggest that the use of zeolite catalysts with large surface area and mesoporous such as ZSM-5, zeolite Y and zeolite A introduced good catalytic activity [[26], [36]]. However, neither metal oxide nor zeolite catalysts can control the oxygen involved in POM reaction.

The disadvantages can be overcome by the use of membranes which are selectively oxygen ion-conducting agents, in which one of the materials that can be used is perovskite oxide. Oxygen supply for POM comes from perovskite lattice oxygen and the amount of oxygen that can be provided depends on the type of perovskite oxide, its membrane thickness, reaction temperature and pressure [40]. Given the extreme reaction conditions (high temperature and pressure), the usable oxygen ion-conducting membrane needs to include the criteria as shown in Figure 1

2
One of the perovskite oxide membranes which has a general arrangement of ABO_3 and high oxygen permeation is Co-based as $\text{LaCoO}_{3-\delta}$ which was first studied by Teraoka et al [37]. In an effort to increase the conductivity of ions and electrons, each side

of A can be substituted ions Ba^{2+} and/or Sr^{2+} while the B side is Fe^{n+} and/or Co^{4+} . However, both Ba^{2+} and Sr^{2+} easily react with CO_2 to form BaCO_3 and SrCO_3 thereby potentially producing a high carbon deposition [[41], [42]]. The substitution of Co^{4+} by Fe^{3+} as in $\text{SrCo}_{0.8}\text{Fe}_{0.2}\text{O}_{3\delta}$ produces membrane performance with oxygen permeation of $3.1 \text{ mL}\cdot\text{cm}^{-2}\cdot\text{min}^{-1}$ at $850 \text{ }^\circ\text{C}$ but low structural and mechanical stability [40]. On the contrary, the substitution of Sr^{2+} with other metal ions such as La^{3+} has a higher valence that uneasily reacts with CO_2 and improves structural stability but still low oxygen flux [[43], [44]].

In addition to the LSCF ($\text{La}_{1-x}\text{Sr}_x\text{Co}_{1-y}\text{Fe}_y\text{O}_3$) group, LSM ($\text{La}_{1-x}\text{Sr}_x\text{MnO}_3$) is also extensively studied for POM reaction due to both structure materials are able to initiate the lattice oxygen to active the CH_4 [[45]]. The oxygen ion conductivity of LSM membrane at $900 \text{ }^\circ\text{C}$ is $10^{-7} \text{ S}\cdot\text{cm}^{-1}$, which is lower than LSCF $\sim 0.2 \text{ S}\cdot\text{cm}^{-1}$. However, LSM has higher reduction activity of O_2 with an electron conductor of $300 \text{ S}\cdot\text{cm}^{-1}$ than LSCF of $230 \text{ S}\cdot\text{cm}^{-1}$ [46]. Therefore, many researchers develop LSCF and LSM-based perovskite oxides, particularly $\text{La}_{0.7}\text{Sr}_{0.3}\text{Co}_{0.2}\text{Fe}_{0.8}\text{O}_3$ and $\text{La}_{0.7}\text{Sr}_{0.3}\text{MnO}_3$ as the potential catalyst for methane conversion, primarily on syngas production.

Conclusion

Partial oxidation of methane (POM) reaction has been studied as a favorable reaction to produce syngas (CO+H₂). However, there were several challenges in the reaction condition such as catalyst usage, effective temperature and pressure to maintain the controlled oxygen reaction with methane. It is due to the high stability of carbon and hydrogen bonding in methane which easily initiates the CO₂ and carbon deposition formation while the presence of oxygen (O₂) in the methane conversion reaction. The heterogeneous catalyst has been developed to solve the problems and contributes significant promising results, such as perovskite oxide which is based on the presence of oxygen vacancy in the crystal structure. The studies indicate the different content on the ionic metal in sites A and B of perovskite revealed significant effects on the performance, especially ionic and electronic conductivity at different temperatures. Nevertheless, a low operating temperature and optimal condition in modified perovskite oxide and the use of other catalysts need to be approached to enhance the commercialization.

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